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2. Balance Everything But O And H. 3. Balance O By Adding H2O (aqueous). 4. Balance H By Adding H+ (assume Acidic). 5. Balance Charge By Adding E-. 6. Balance E- By Multiplying By A Common Factor. 7. Add Together And Simplify. Balancing In Base: Balance As If In Acid; Then Add As Many OH- Ions To Both Sides As There Are H+ Ions. 14th, 2024

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____ 38. A Compound Contains Carbon, Oxygen, And Hydrogen. Analysis Of A Sample Showed That It Contained By Mass 68.9% Carbon And 4.92% Hydrogen. What Is The Simplest Formula For This Compound? A. C6H6O2 D. C6H4O3 B. C7H6O2 E.

C7H8O C. C8H6O2 ____ 39. A Compound Is Known To Contain Only 3th, 2024

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