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14) Pure Acetic Acid ($\text{HC}_2\text{H}_3\text{O}_2$) is a liquid and is known as glacial acetic acid. Calculate the molarity of a solution prepared by dissolving 15.00 mL of glacial acetic acid at 25°C in sufficient water to give 500.0 mL of solution. The density of glacial acetic acid at 25°C is 1.049 g/mL. A) 1.89 $\times 10^{-3}$ B) 31.5 C) 0.525 D) 5.25 $\times 10^{-4}$ E) 0.0315 Feb 7th, 2024

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Question. 1) ΔH For The Reaction $5 \text{ F}_2 (\text{g}) + 3 \text{ IF}_3 (\text{g}) \rightarrow 5 \text{ F}_2 (\text{g}) + 3 \text{ IF}_3 (\text{g})$ Is _____ KJ, Give T Jun 7th, 2024

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