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Might Be Easiest To Balance The Equation With Mostly
Whole Numbers: $2 \text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow 2 \text{NO}_2 + 3 \text{H}_2\text{O}$.
The Question Asks About The Amount Of Oxygen
Reacting With ONE Mole Of Ammonia, Thus Cut The $\frac{7}{2}$
(3.5) Of Oxygen In Half To 1.75 2. C Balance: CH Apr
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Unit 10 Stoichiometry Test Answers

Unit 10 Stoichiometry Test Answers Monday, February 13

Agenda: 1) Go Over Quest 2) Introduce

Stoichiometry-S'more Activity HW: S'More Worksheet

#2 Tuesday, February 14

Agenda: Lab -Introduction To Stoichiometry

HW: 1) Lab Calculations 2) Read P

237-240 Wednesday, February 15

Agenda: 1) Go Over Lab 2) Worksheet #3

HW: 1) Read P 242 2) Video Notes Thursday, Mar 9th, 2024

Stoichiometry Test Review March 21/22

5. Given The Incomplete Equation: $2 \text{N}_2\text{O}_5(\text{g}) \rightarrow$ Which Set Of Products Completes And Balances The

Incomplete Equation? 6. Base Your Answer To The

Following Question On Given The Balanced Equation

Representing A Reaction: $\text{C}_3\text{H}_8(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) +$

$4\text{H}_2\text{O}(\text{g})$ What Is The Total Number Of Moles Of $\text{O}_2(\text{g})$

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