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Quantitative Relationships Between The Amounts Of

Reactants Used And Amounts Of Products Formed By A

Chemical Reaction Is Called Stoichiometry.

Stoichiometry Is Based On The Law Of Conservation Of

Mass. Recall From Chapter 3 That The Law States That

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Aqueous Solution ... Buildup Of Partial Negative Charge
Near The Oxygen End Of The Molecule And Partial
Positive Charge Near The Hydrogen Atoms. Thus, The

Negative End Of A Water Molecule Is Strongly

Attracted To ... Feb 1th, 2024 Chapter 3

Stoichiometry Chapter 3 Stoichiometry 3-7 EXAMPLE

PROBLEM: Determine The Molar Mass Of A Compound

Calculate The Molar Mass For Each Of The Following

Compounds: (a) 2-Propanol, $\text{CH}_3\text{CH}(\text{OH})\text{CH}_3$ (b) Iron(II)

Phosphate, $\text{Fe}_3(\text{PO}_4)_2$ SOLUTION: You Are Asked To

Calculate The Molar Mass Of A Compound. You Are

Given The Compound's Formula. Jan 7th, 2024.

CHAPTER 2: Atoms, Molecules And Stoichiometry 2.5

Stoichiometry And Equations Stoichiometry 1)

Stoichiometry Is The Proportion Of Things Either

Reacting Or Combining. 2) In Compounds, It Refers To

The Ratio In Which The Atoms Are Combined Together.

For Example, Water, H_2O Has A Stoichiometry Of 2

Hydrogen To 1 Oxygen. 3) It Also Refers To The

Reacting Proportions In A Chemical Equation. Apr 9th,

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