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CHAPTER 15 Acid-Base Titration And PH "molar Hydronium Ion Concentration." In Pure Water At 25°C, $[H_3O^+] = 1.0 \times 10^{-7} M$ And $[OH^-] = 1.0 \times 10^{-7} M$. Two Water Molecules React To Form One Hydronium Ion And One Hydroxide Ion. Therefore, When Pure Water Self-Ionizes, There Should Be An Equal Number Of Hydronium Ions And Hydroxide Ions. Mar 24th, 2024.

CHAPTER 16 Acid-Base Titration And PH Neutral, Acidic, And Basic Solutions Because The Hydronium Ion And Hydroxide Ion Concentrations Are The Same In Pure Water, It Is Neutral. Any Solution In Which $[H_3O^+] = [OH^-]$ Is Also Neutral. Recall From Chapter 15 That Acids Increase The Concentration Of H_3O^+ in Aqueous Solutions, As Shown In Figure 16- Mar 4th, 2024

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TABLE OF CONJUGATE ACID-BASE PAIRS Acid Base K_a (25 C) TABLE OF CONJUGATE ACID-BASE PAIRS Acid Base K_a (25 OC) $HClO_4$ ClO_4^- - H_2SO_4 HSO_4^- - HCl Cl^- - HNO_3 NO_3^- - H_3O^+ H_2O H_2CrO_4 $HCrO_4^-$ - 1.8×10^{-1} $H_2C_2O_4$ (oxalic Acid) Mar 22th, 2024

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