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D) N₂O₅ 25.What Is The Empirical Formula Of A
Compound That Contains 30.4% Nitrogen And 69.6%
Oxygen By Mass? A) SO B) SO₂ C)SO₃ D) SO₄ 26.A
Compound Consists Of 40.% Sulfur And 60.% Oxygen
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Ca(NO₃)₂ ... Use The Following Balanced Equation.

Identify The Limiting Reactant When 1.150 Grams Of HgO React With 12.46 Grams Of Cl₂. Convert To Mol
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Chemistry 11 Stoichiometry Review Worksheet **Name: KEY**

Chemistry 11 Stoichiometry Review Worksheet

Name: KEY 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1\text{mol}$

$122.55\text{g} \times 2\text{O}_2 = 1\text{KClO}_4 \times 32.00\text{g} = 3.11\text{g O}_2$

%Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%? pure FeCO_3 6th, 2024

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In A Camping Stove? 2. How Many Moles Of O₂

Molecules Should Be Supplied To Burn 1 Mol Of CH₄ ...

Was At One Time Considered For Use As A Rocket Fuel.

... How Many Formula Units Of Calcium Oxide (CaO)

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 $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
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