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12 1 Stoichiometry Worksheet Read PDF Chapter 12 1 Stoichiometry Worksheet Answers Figure 11.4.1 Is Shown In Figure 12.2.1. We Can Use The Balanced Chemical Equation For The Reaction And Either The Masses Of Solid Reactants And Products May 1th, 2024.

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Stoichiometry Worksheet #1 Answers  
1. Given The Following Equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ , Show What The Following Molar Ratios Should Be.  
A.  $\text{C}_4\text{H}_{10} / \text{O}_2$  B.  $\text{O}_2 / \text{CO}_2$  C.  $\text{O}_2 / \text{H}_2\text{O}$  D.  $\text{C}_4\text{H}_{10} / \text{CO}_2$  E.  $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$   
2. Given The Following Equation:  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$   
A. How Many Moles Of  $\text{O}_2$  Can Be Produced By ...  
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33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$  Actual yield = 1.45g Theoretical yield = 5.95g  
 $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times 2 \text{ O}_2$   
 $1 \text{ KClO}_4 \times 32.00 \text{ g } 1 \text{ mol } = 3.11 \text{ g O}_2$  %Yield =  $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$   
34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$  Percent purity = 62.8%  
g pure  $\text{FeCO}_3$  Jan 2th, 2024  
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