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 $RT [\text{Red}] E = E_{\text{O}} - \frac{RT}{nF} \ln \frac{[\text{Ox}]}{[\text{Red}]}$ Where R = Gas Constant T = Absolute Temperature n = Number Of Electrons In Reaction F = Faraday's Constant E = Potential E_{O} = Standard Potential $[\text{Red}]$ = Molar Concentration Of Reduced Form Of Specie
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