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$\Delta G = -nFE$  (2.1.36) The Free Energy Of The Cell Reaction Is Given By,  $\Delta G = \Delta G^0 + RT \ln \frac{a_{\text{R}}}{a_{\text{O}}}$  (2.1.38) Which Since  $\Delta G = -nFE_{\text{rxn}}$  And  $\frac{a_{\text{H}^+}}{a_{\text{H}_2}} = 1$  Becomes,  $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{R}}}{a_{\text{O}}}$ . (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Jan 12th, 2024

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