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Charles, Are Summarized In The Ideal Gas Law:  $PV =$   
 $nRT$   $P =$  Pressure  $V =$  Volume  $n =$  Moles Of Gas,  $R =$   
Universal Gas Constant  $T =$  Temperature. The Value Of  
 $R$  Varies With The Mar 2th, 2024 IDEAL 2360 IDEAL  
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1. The He Gas Because It Is Less Massive And Moving With A Higher Average Velocity

Correct 2. They Are The Same Since The Pressure Is The Same 3. The O<sub>2</sub> Since It Has A Higher Average Momentum Since It Is More Massive 4. Jun 2th, 2024.

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