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 1.1 Rate Of Reaction And Dependence On Temperature We Will Once Again Look At
 The Formation Of Ammonia (NH₃) From Nitrogen And Hydrogen (see Section
 Chemical Equilibrium Of The Thermodynamics Chapter). This Reaction Follows The
 Equation: N₂ + 3H₂ → 2NH₃ (1) H₀ = 92 KJ Mol S₀ = 192 J Mol K^{1st}, 2024Wiley
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 THERMODYNAMIC PROPERTIES OF CHEMICAL SUBSTANCES
 Na Sodium
 Tetrafluoroaluminate -1869.0 -1827.5 345.7 105.9 AlH Aluminum Hydride 259.2
 231.2 187.9 29.4 AlH₃ Aluminum Hydride -46.0 30.0 40.2 AlH₄K Potassium
 Aluminum Hydride -183.7 AlH₄Li Lithium Aluminum Hydride -116.3 -44.7 78.7 83.2
 AlH₄Na Sodium Aluminum Hydride -15.5 All Aluminum Monoiodide 65.5 36.0 2th,
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J/mol K C P J/mol K F H ° KJ/mol F G ° KJ/mol S ° J/mol K C P J/mol K F H ° KJ/mol F G °
KJ/mol S ° J/mol K C P J/mol K Substances Not Containing Carbon: Ac Actinium 0.0
56.5 27.2 406.0 366.0 188.1 20.8 Ag Silver 0.0 42.6 ... 2th, 2024Thermodynamic
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Calculations Free Energy Calculations Are Carried Out In Much The Same As
Enthalpycalculations. There Are Tables Of Standard Free Energies Of Formation Of
Compounds.Elements In Their Standard States Areassigned Zeroastheir Gibbs Free
Energy Of Formation. Consider The Following Chemical Reaction: $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g})$
 $\rightarrow 2\text{HCl}(\text{g})$. 1th, 2024.

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Formation And Energy Of HOMO. The Second PC Is Strongly Influenced By Energy Of
LUMO, While Substitution Pattern Parameters, Number Of Chlorine Atoms At α -
positions And Vicinal (adjac-ent) Carbon Atoms 1th, 2024Selected Values Of
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Valueof ΔH_f° represents the apparent Enthalpy Of the Reaction Of Formation Of the

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