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# A: Example Problem: Batch Chemical Reactor

For This Example, Two Process Drawings Are Presented: Exhibit Al: Process Flow Diagram (PFD) With A Material Balance And Exhibit A2: Piping & Instrumentation Diagram (PMD). Physical And Hazardous Properties Were Obtained From Open Technical Lit- Erature And Company Files. The Heat Of Reaction And Runaway Potential Data Jan 1th, 2024

## **Fundamentals Of Chemical Reactor Theory1**

Ea = Activation Energy [J/mol, Cal/mol] R = Gas Constant =  $8.314 \text{ J/mol} \cdot \text{K} = 1.987 \text{ Cal/mol} \cdot \text{K} T = \text{Absolute Temperature}$  [K] The Mass Balance Mass Is A Conservative EntityIV, Hence Given A Control Volume V The Sum Of Mass Flows Entering The System Will Equal The Sum Exiting Minus Apr 2th, 2024

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1 Chemical Reactions 1.1 Rate Of Reaction And Dependence On Temperature We Will Once Again Look At The Formation Of Ammonia (NH 3) From Nitrogen And Hydrogen (see Section Chemical Equilibrium Of The Thermodynamics Chapter). This Reaction Follows The Equation: N 2+3H 2 2NH 3 (1) H0 = 92 KJ Mol S0 = 192 J Mol K Feb 1th. 2024

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