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 $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$  C. Jan 12th, 2024  
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KEY Chemistry 11 Stoichiometry Review Worksheet Name: \_\_\_KEY\_\_\_ 33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$  Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times 2 \text{ O}_2$   
 $1 \text{ KClO}_4 \times 32.00 \text{ g } 1 \text{ mol} = 3.11 \text{ g O}_2$  % Yield =  $\frac{1.45 \text{ g } 3.11 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$  34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$  Percent purity = 62.8%? g pure  $\text{FeCO}_3$  Apr 6th,  
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Stove? 2. How Many Moles Of  $\text{O}_2$  Molecules Should Be Supplied To Burn 1 Mol Of  
 $\text{CH}_4$  ... Was At One Time Considered For Use As A Rocket Fuel. ... How Many  
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