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1. A Solution Is Prepared By Mixing 100.0 G Of Water, H_2O , And 100.0 G Of Ethanol, C_2H_5OH . Determine The Mole Fractions Of Each Substance. 2. The Molality Of An Aqueous Solution Of Sugar ($C_{12}H_{22}O_{11}$) Is 1.62m. Calculate The Mole Fractions Of Sugar And Water. 3. Calculate The Molality Of 25.0 Grams Of 3th, 2024

Mole Fraction Molality Molarity - University Of Texas At ...

Mole Fraction Sugar Of Solution $\chi_{\text{Sugar}} = (0.292 \text{ Mol}) / (0.292 + 55.6) = .00522$ Mole Fraction Water Of Solution $\chi_{\text{Water}} = (55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$ (or $1 - \chi_{\text{Sugar}}$) Molality $M = (.292 \text{ Mol}) / (1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$ Molarity $M = (.292 \text{ Mol}) / (1 \text{ L}) = 0.292 \text{ Mol L}^{-1}$ 3th, 2024

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3.mole Fraction = 0.950 4.mole Fraction = 0.0360 5.mole Fraction = 0.925 Nsci140802exam 018 10.0points 3.3 G Of NaCl (58.4 G/mol) And 9 G Of KBr (119.8 G/mol) Were Dissolved In 40 G Of Water (18.0 G/mol). What Is The Mole Fraction Of KBr In The Solution? Nsci140805 019 10.0points How Many G 2th, 2024

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As $1 \text{ Atm} = 1.013 \times 10^5 \text{ Pa}$, The Osmotic Pressure Is $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$ Hence $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ J K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$ As $1 \text{ M}^3 = 1000 \text{ L}$, This Corresponds To $C = 1.5 \times 10^{-4} \text{ Mol L}^{-1}$ 1th, 2024

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