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2. What Volume Of 2.50M KCl Is Needed To Make 2.00 L Of A 1.00M Solution?  $M_1V_1 = M_2V_2$   
 $2.50 \times 1.00 = 1.00 \times V_2$   
 $V_2 = 2.50$  L  
3. Concentrated  $H_2SO_4$  Is 18.0M What Volume Is Needed To Make 2.00 L Of 1.00M Solution?  $18.0 \times V_1 = 1.00 \times 2.00$   
 $V_1 = 0.111$  L  
4. Concentrated HCl Is 12.0M What 1th, 2024

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