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10. Write The Molecular And Net Ionic Equations For The Following Aqueous Reactions. If No Reaction Occurs, Write NR. A) $\text{Ba}(\text{NO}_3)_2 + \text{K}_2\text{SO}_4$ B) $\text{Mg}(\text{NO}_3)_2 + \text{K}_2\text{SO}_4$ C) $\text{CaCl}_2 + \text{Al}(\text{NO}_3)_3$ D) $\text{NiCl}_2 + \text{NaBr}$ E) $\text{Na}_2\text{CO}_3 + \text{MgBr}_2$. 11. Write The Net Ionic Equation For Each Of The Following. Use The General Rules For Solubility (Bauer Table 5.3, P 177). Apr 14th, 2024

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