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Therefore The Ratio Is 1:1, So The Empirical Formula = HF B. $(4.267g \text{ UF}_6 \times 1 \text{ Mol UF}_6 \times 6 \text{ Mol F}) / (352.03g \times 1 \text{ mol}) = 0.07273 \text{ Mol F}$ Original Compound Gas: $(0.0485 \text{ Mol F}) / (0.07273 \text{ Mol F}) = 66.7\%$ In Gas Or 922/1382 Solid: 33.3% In Solid Or 460/1382 C. $(4.267g \text{ UF}_6 \times 1 \text{ Mol UF}_6 \times 1 \text{ 4th, 2024}$

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