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Answers: Moles And Stoichiometry Practice Problems 1) How Many Moles Of Sodium Atoms Correspond To 1.56×10^{21} Atoms Of Sodium? $1.56 \times 10^{21} \text{ Atoms Na} \times \frac{1 \text{ Mol Na}}{6.022 \times 10^{23} \text{ Atoms Na}} = 2.59 \times 10^{-3} \text{ Mol Na}$ 2) Determine The Mass In Grams Of Each Of The Following: A. $1.35 \text{ Mol Fe} \times 55.845 \text{ g/mol Fe} = 75.4 \text{ g Fe}$ 1 Mol Fe B. 24.5 Mol O Mar 17th, 2024

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