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Pressure Of 2.00 Bar And A Temperature Of 273 K Is Taken To A Final Pressure Of 4.00 Bar By The Reversible Path Defined By $P/V_m = C$, Where C Is A Constant And V_m Is The Molar Volume. (a) Show That C Is 0.176 Bar Mol L⁻¹. (b) Calculate The Final Temperature For This Process. (c) Calculate T 22th, 2024.

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