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Table C.2 (K) Nitrogen ( N <sub>2</sub> )	
100	150
200	250
300	350
450	550
700	1,000
1,100	1,200
1,300	

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 87 The Total Number Of Moles, N, Is The Sum Of The Number Of Moles Of All The Components Of The Mixture:  $N = N_1 + N_2 + \dots$  It Follows That The Sum Of All The Mole Fractions Of The Mixture Must Also Equal 1.  $X_1 + X_2 + \dots = 1$  The Mass Of Component I In A Mixture Is The Product Of The Number Of Moles Of I And Its Molecular Weight,  $M_i$ . The Mass Of The Mixture Is Therefore The Sum,  $M = N_1M_1 + N_2M_2 + \dots$  May 9th, 2024.

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